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the equation: $\text{Sn}(s) + 2\text{HF} \dots$

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Determine whether each of the
following is an example of observation

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Answers, a theory, a hypothesis, a

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Questions in the space provided. 1.
Complete the following conversions: a.
100 mL 0.1 L b. 0.25 g 25 cg c. 400
cm³ 0.4 L d. 400 cm 3 0.0004 m 2. For
each measuring device shown below,
identify the quantity measured and tell
when it would remain constant ...

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2 Measurements and Calculations

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Objectives 12.2.1 Construct mole ratios from balanced chemical equations and apply these ratios in stoichiometric calculations. 12.2.2 Calculate stoichiometric quantities from balanced chemical equations

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Answers using units of moles, mass,
representative particles, and volumes
of gases at STP. Guide for Reading

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Answers to Chapter 122 ...

chapter-22-accounting-changes-and-error-analysis-multiple ...

Answers 1. a. $2 \text{ Al}_2\text{O}_3 \rightarrow 4 \text{ Al} + 3 \text{ O}_2$ or $2 \text{ Al}_2\text{O}_3 \rightarrow 4 \text{ Al} + 3 \text{ O}_2$
 $2 \text{ Al}_2\text{O}_3$ moles $\text{Al}_2\text{O}_3 = (2.3)(2)/4 = 1.2$
moles b. Using the same ratios, moles
 $\text{O}_2 = (3.9)(3)/2 = 5.6$ moles 2. a. 2

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moles Fe gives 3 moles H₂, moles H₂
 $2 = (1.7)(3)/2 = 2.6$ moles b. 3 moles H₂
2SO₄ gives 1 mole product moles
yield = 3 x 2.8 = 8.4 moles 3. Mole
ratios: 2 mol Mg/ 2 mol MgO = 1 mol
Mg: 1 mol product 1 mol O

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