

# Read Book Chapter 10 Chemical Quantities Study Guide Answer Key

## Chapter 10 Chemical Quantities Study Guide Answer Key

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(Empirical/Molecular Formula)

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Chapter 10 - Gases Relationship among n, m, M, N  
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~~Quantities?~~ Significant Figures - A Fast Review!  
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Chapter 10 Chemical Quantities Study  
Chemistry Chapter 10 Chemical Quantities. STUDY.

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PLAY. Mole. (mol) of a substance is  $6.02 \times 10^{23}$  representative particles of that substance and is the SI unit for measuring the amount of a substance.

Avogadro's Number. the number of representative particles in a mole,  $6.02 \times 10^{23}$ . Representative Particle.

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Chapter 10 Chemical Quantities. mole. Avogadro's number. representative particle. molar mass. the SI unit for measuring the amount of a substance.

number of representative particles in a mole,  $6.02 \times$

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$10^{23}$ . refers to the species present in a substance: usually atoms, m.... the mass of one mole of a pure substance.

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Chapter 10 Chemical Quantities. mole. Avogadro's number. representative particle. molar mass. the SI unit for measuring the amount of a substance.

number of representative particles in a mole,  $6.02 \times 10^{23}$ . refers to the species present in a substance: usually atoms, m.... the mass of one mole of a pure substance.

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Created by. claire\_kutchka. mrs. wright. Terms in this set (38) standard temperature and pressure. 0 C and 101.3 kPa. the volume occupied by a mole of any gas at STP (22.4 L) molar volume.

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Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_ Chapter 10 -

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Chemical Quantities Avogadro's Number One important property of a mole is that it means a definite number of particles just like a dozen means a number of particles. While a dozen is only 12 particles a mole is a larger number –  $6.02 \times 10^{23}$  particles.

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Ch 10 - Avogadro's Number.docx - Name Date Class  
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Chemistry Section 10 Study Guide Chemical  
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Chapter 10 Chemical Quantities91 SECTION 10.1 THE

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MOLE: A MEASUREMENT OF MATTER (pages 287–296)

This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287–289) 1.

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SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296)

Chemical Quantities THE MOLE AND QUANTIFYING

MATTER 10.1 The Mole: A Measurement of Matter

Essential Understanding The mole represents a large number of very small particles.

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Chemistry Section 10 Study Guide Chemical Quantities Chemistry Section 10 Study Guide Chemical Quantities corresponding mass of another substance participating in the same reaction. For a related section, see Equation Stoichiometry Problems with Mixtures on our Web site. Chapter 10 Chemical Calculations and Chemical Equations Chemistry

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Section 10 Study Guide TEACHER GUIDE AND ANSWERS Chemistry: Matter and Chemistry Section 10 Study Guide Chemical Quantities Study Flashcards On Grade 10 ...

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Chemistry Section 10 Study Guide Chemical Quantities

CHAPTER 10: Chemical Quantities BASICS: □ The basic unit that is used to determine the amount of a chemical substance is called a mole □ A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance □ The mole was founded by a scientist named Avagadro, and he

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