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Kinetic Molecular Theory Key Questions 1. What causes a gas to exert pressure when confined in a container? 2. How does the total volume of gas particles compare to the volume of the space between the gas particles? betwun 15 arc.) 3. As the temperature of a gas decreases, what change occurs in the amount of kinetic energy? Uaeases 4.

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TES OF THE KINETIC MOLECULAR THEORY Gases consist of tiny particles (atoms or molecules), These particles are so small, compared with the

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distances between them that the volume (size) of the individual particles can be assumed to be negligible (zero). The particles are in constant random motion, colliding with the walls of the container.

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